Subs., 0.1330: (Kjeldahl), 7.3 cc. 0.1 *N* HCl.
Calc. for $C_{10}H_{15}O_2N$: N, 7.74. Found: 7.69.

2,4-Diethoxy-chloroacetanilide.—One g. of the base was dissolved in a mixture of 5 cc. of acetic acid and 5 cc. of saturated sodium acetate solution, diluted with 30 cc. of 50% acetic acid, and chloroacetylated in the usual way. After dilution the product was filtered off and recrystallized from 85% alcohol, separating as delicate, woolly needles, which melt at 102–3° with slight preliminary softening. The substance is quite soluble in alcohol at room temperature, very readily in acetone, chloroform, or benzene. It is difficultly soluble in cold ligroin, readily on heating, and also dissolves sparingly in boiling water.

Subs., 0.1541: (Kjeldahl), 5.85 cc. 0.1 *N* HCl.
Calc. for $C_{12}H_{16}O_3NCl$: N, 5.44. Found: 5.32.

NEW YORK CITY.

NOTE.

Correction.—In the paper on the "Constitution of Capsaicin," appearing in the July number,[1] vanillyl amine is by error described as 3-hydroxy-4-methoxy-benzylamine on pp. 1118, 1119, 1121, instead of 4-hydroxy-3-metnoxy-benzylamine. E. K. NELSON.

NEW BOOKS.

An Advanced Course in Quantitative Analysis. By HENRY FAY, PH.D., D.Sc. John Wiley & Sons, Inc., New York, 1917. Pp. vi + 111.

Text books on quantitative analysis may usually be placed in one of two groups; either the author has selected methods with reference primarily to the training of the student in the principles of quantitative analysis and in stoichiometry and has ignored the practicability of the methods described or else he has based his text book upon current and so-called "standard methods." In the latter case, too often, simplified manipulation, rapidity and occasionally great accuracy are the reasons for the selection of the material and the question of the effect upon the student of such a course is apt to be overlooked; the main idea being to have the student become proficient in the methods of analysis used in the industries.

Professor Fay has tried to produce a text which falls in neither extreme group. The material he has selected has been "chosen to illustrate principles and to train the student in manipulation." Since the author's work and interests lie mainly in iron, steel and commerical alloys, in which field he is a recognized authority, it is not surprising to find the bulk of the book devoted to the analysis of steel and closely associated raw and finished materials. Although the methods given have been selected

[1] THIS JOURNAL, 41, 1115 (1919).